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1) calculate moles of each element (if given mass) 2) divide each mole quantity by the smallest mole value from step 1. 3) use whole number values from step 2 as the subscripts to the original equation to write the empirical formula. Percent Composition. -the relative amounts of each element in a compound.

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2nd Semester Chemistry-575 Final Exam Review Multiple Choice Identify the choice that best completes the statement or answers the question. ____ 1. How many moles of tungsten atoms are in 4.8×10^{25} atoms of tungsten? a. 8.0×10^2 moles c. 1.3×10^{11} moles b. 8.0×10^1 moles d. 1.3×10^{12} moles

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