

Chemistry 2 Study Guide

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~~Intro to Chemistry, Basic Concepts - Periodic Table, Elements, Metric System \u0026 Unit ConversionChemistry 2 Study~~
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~~Chemistry Unit 2 Study Guide 1. All matter is made of indivisible, indestructible atoms 2. Atoms of a given element are~~
~~identical in their physical and chemical properties 3. Atoms of different elements have different physical and chemical~~
~~properties 4. Atoms of different elements combine in simple, ...~~

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~~You'll need to know this as you study Chemistry II. It has the form: In this equation ΔG° is the Gibbs Free Energy of a~~
~~reaction under standard conditions of 1 atm (or 1 bar) for gases and 1 M for solutions at 25°C; ΔH° is the enthalpy of the~~
~~reaction under standard conditions; T is the Kelvin temperature; and ΔS° is the entropy of the reaction under standard~~
~~conditions.~~

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~~Reaction: Reactant- a substance converted into another substance during a chemical reaction. Product- a substance formed~~
~~from another substance or substances during a chemical reaction. o Products are different from reactants in the~~
~~arrangement of their component atoms.~~

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~~Study Guide for Chemistry 2 Lab Safety 1. What is correct lab attire? A. Closed toe shoes with socks B. Apron or lab coat C.~~
~~Safety Goggles D. All the above 2. You can bring food and drinks into the lab. A. True B. False 3. What is a vent hood?-A~~
~~vent hood is the glass enclosure that pulls vapors out of the air and filters them. 4.~~

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~~Basic Chemistry. In the learn based tool, prompt with term and copy the "whole" answer and paste it into the answer box.~~

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~~2 2 Atomic charge is determined by the relative number of electrons and protons of an atom. While the number of protons~~

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of an atom is stable, the number of electrons varies. Atoms with a non-zero charge are called ions. An atom with an equal number of protons and electrons has a neutral charge.

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List the 4 steps of the scientific method and explain what they do/ why they are important. observation- qualitative and quantitative data. hypothesis- tentative explanation to be tested. experiments- a set of controlled observations that test the hypothesis. conclusion- a statement that confirms or denies the hypothesis based off the information from the experiment.

Chemistry 1-2 STUDY GUIDE Flashcards | Quizlet

Chemistry Exam 2 Study Guide. Lecture 12: February 10th. Scientific notation: a compact way to write very large or very small numbers. Units o Specific units of measurement are defined by the International System of Units o Ex. Meters for length/distance, kilograms for mass, kelvin for temperature; Units and Equivalentents

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Unit 2 Study Guide Chemistry. STUDY. PLAY. Democritus. Greek philosopher who formulated the atomic theory of the universe. Democritus's theory and model-atoms were small-same material-different shapes and sizes -infinite and indivisible. Dalton's theory-all elements are composed of atoms

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